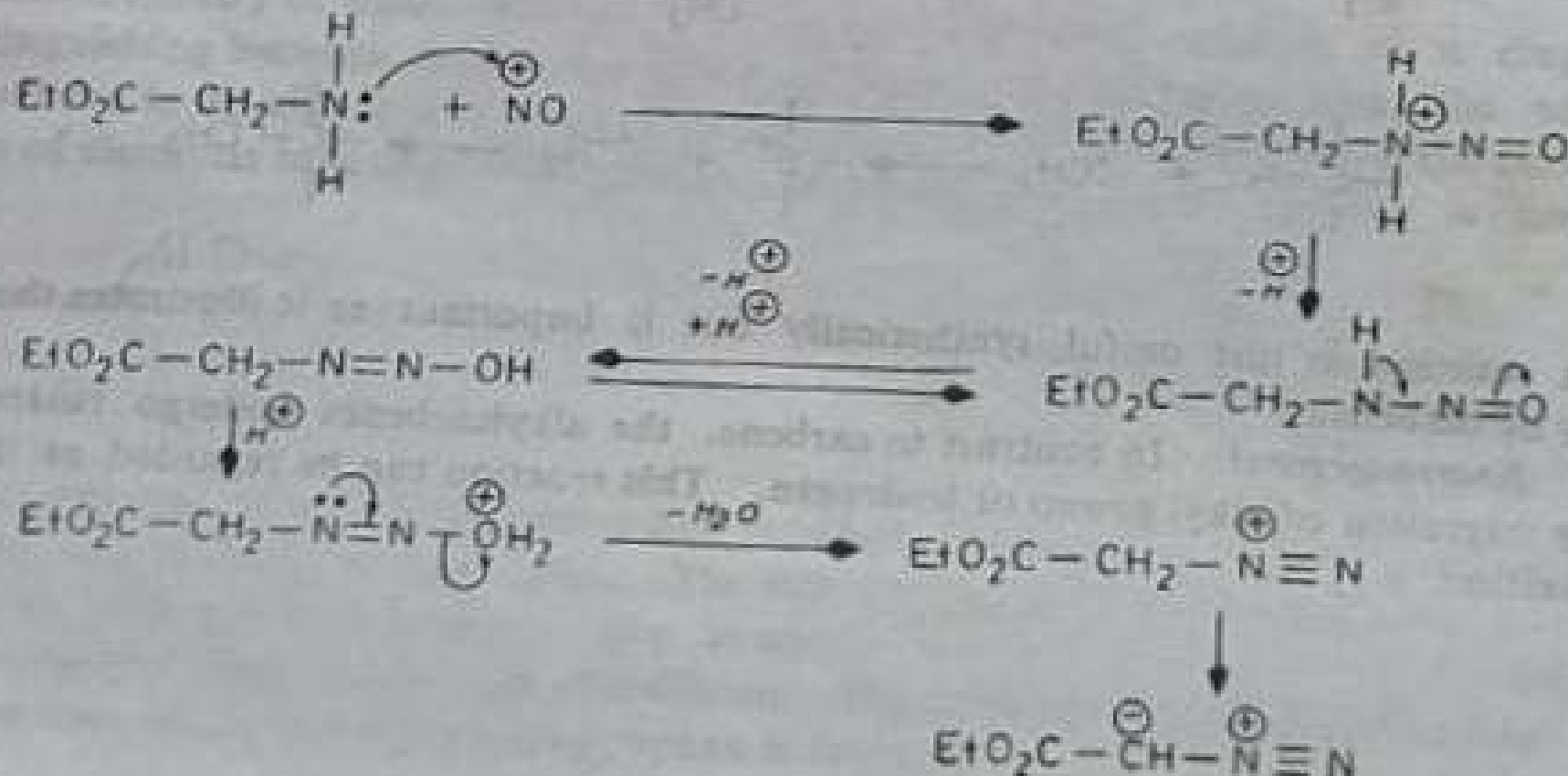
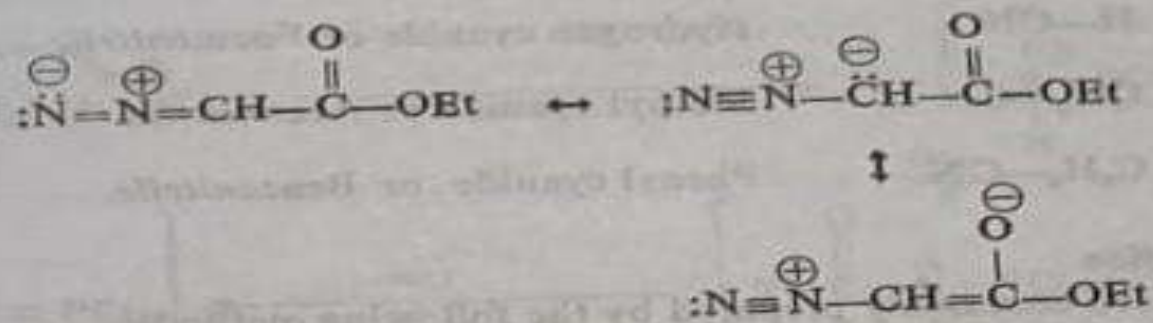


19.4.2 Diazoacetic Ester

It is prepared by treating a cold solution of the hydrochloride of ethyl aminoacetate with cold aqueous sodium nitrite solution. The reaction takes place by the initial attack of nitrosonium ion on the amino group forming a diazonium cation which loses a proton to water forming the diazo compound.



It is a yellow oil (b.p. 140/720 mm) insoluble in water but soluble in organic solvents. It is quite stable, so much so that it can be distilled. The stability is attributed to extended conjugation.



The chemical nature of diazoacetic ester is quite similar to that of diazomethane.

